**Chemical Compound Naming & Formula Writing Practice**

**Naming Ionic Compounds**

1. NaCl \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
2. CaCl2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
3. MgO \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
4. Li2S \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_­­\_\_
5. KI \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_­­­\_\_\_
6. BeF2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
7. Sr3P2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
8. HCl \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
9. KF \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
10. Na2O \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**Writing Ionic Formulas**

1. Calcium oxide \_\_\_\_\_\_\_\_\_\_\_
2. Beryilium chloride \_\_\_\_\_\_\_\_\_
3. Sodium sulfide \_\_\_\_\_\_\_\_\_\_\_
4. Magnesium phosphide \_\_\_\_\_
5. Potassium fluoride \_\_\_\_\_\_\_\_
6. Calcium fluoride \_\_\_\_\_\_\_\_\_\_
7. Lithium oxide \_\_\_\_\_\_\_\_\_\_\_\_
8. Hydrogen fluoride \_\_\_\_\_\_\_\_\_
9. Rubidium oxide \_\_\_\_\_\_\_\_\_\_\_
10. Barium sulfide \_\_\_\_\_\_\_\_\_\_\_

**Naming Covalent Compounds**

1. SO3 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
2. N2S \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
3. PH3 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
4. BF3 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
5. P2Br4 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
6. CO \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
7. SiO2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
8. SF6 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
9. NH3 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
10. NO2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**Writing Covalent Formulas**

1. Nitrogen tricholoride \_\_\_\_\_\_\_
2. Boron monocarbide\_\_\_\_\_\_\_\_
3. Dinitrogen trioxide \_\_\_\_\_\_\_\_\_
4. Phosphorous pentafluoride\_\_\_\_
5. Sulfur dibromide \_\_\_\_\_\_\_\_\_\_\_
6. Oxygen difluoride \_\_\_\_\_\_\_\_\_\_
7. Diboron tetrahydride \_\_\_\_\_\_\_\_
8. Carbon disulfide \_\_\_\_\_\_\_\_\_\_\_
9. Nitrogen monoxide \_\_\_\_\_\_\_\_\_
10. Dinitrogen decoxide\_\_\_\_\_\_\_\_\_

**Mixed Practice – Determine if Ionic or covalent and then name appropriately.**

1. S2O4 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
2. AlCl3 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
3. CaF2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
4. C3O10 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
5. KBr \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
6. MgF2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
7. CO2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
8. H2S \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
9. H3S \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
10. Ba3N2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**Mixed Practice – Determine if Ionic or covalent and then write correct formula.**

1. Sulfur trioxide\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
2. Aluminum oxide \_\_\_\_\_\_\_\_\_\_\_\_\_
3. Potassium chloride \_\_\_\_\_\_\_\_\_\_
4. Calcium sulfide \_\_\_\_\_\_\_\_\_\_\_\_\_
5. Hydrogen triphosphide \_\_\_\_\_\_\_
6. Nitrogen heptafluoride \_\_\_\_\_\_\_\_
7. Triphosphorous hexabromide\_\_
8. Boron fluoride \_\_\_\_\_\_\_\_\_\_\_\_\_\_
9. Cesium fluoride \_\_\_\_\_\_\_\_\_\_\_\_\_
10. Bromine monosulfide \_\_\_\_\_\_\_\_