**Know the definitions for the following:**Solvent – liquid (usually water) that dissolves the solute to make the solution Solute – solid (usually) that is added to solute and dissolves to make solution Solution – homogeneous mixture of solvent and solute
Concentration – a measure of how much solute per volume of solvent
Molarity – concentration; moles of solute/liters of solvent

Arrhenius Acid and Base – acid donates H+ and base doneates OH-

Bronsted Lowry Acid and Base – acid donates H+ and a base accepts H+

Neutralization – acid + base 🡪 salt + water

Titration – process used to determine concentration of an acid using neutralization

**Solutions/Solubility:**What factors will increase the rate of solubility? Increasing temperature; breaking into pieces; stirring

**Calculating Concentration** (M = moles/liters)

What is the molarity of 50 grams of NaCl in a 100ml of water?

Convert grams 🡪 moles: **50 grams NaCl / 58.5 grams (from periodic table) = .85 moles** Divide moles by volume 🡪 **.85 moles / 0.1 L = 8.5 M**

**Dilutions (M1V1 = M2V2)**

If I have 500 mL of a 0.5M NaCl solution, what will the concentration be if the final volume of the solution is 600 mL? **(500ml)\*(0.5M) = (600ml)( ? M) = 0.42 M**

How many liters of 0.5M HCl stock solution would you need to prepare 2 liters of 0.25 M HCl solution? **(2 L)\*(0.25 M) = (? L) \* (0.5 M) = 1 M**

**Acids and Bases:**

Know the properties of acids and bases:



1. Write definitions for, or explain the meaning of: Bronsted-Lowry acid and base, neutral, strong acid, strong base, weak acid, weak base, concentrated, dilute, ionization, dissociation
2. Know general acid base reactions:



1. Be able to recognize acids, bases and salts from their chemical formulas.
2. What number do you get when you add up the pH and pOH of a solution?

pH + pOH = 14

1. What is pH?

pH is a numerical scale used to express the acidity of a solution.

1. What is the range of pH for an acid?  For a base?  For a neutral compound?

0-7 is acidic, 7 is neutral, 7-14 is basic

1. What happens to pH of a solution as you increase the hydroxide (OH-) ion concentration?

As you increase OH- concentration, the pH of a solution will increase.

1. What happens to the pH of a solution as you increase the hydronium (H+) ion concentration?

As you increase the H+ concentration of a solution, the pH will decrease.

1. Classify the following solutions from most acidic to lease acidic based on their pH and then classify the solutions as most basic to most acidic based on their pH:

Most acidic to most basic: 2, 5, 8, 10, 13

Most basic to most acidic: 13, 10, 8, 5, 2

1. What do you use litmus (pH paper) for?  If it turns red, what does it indicate about a solution?  If it turns blue, what does it indicate about a solution?

You use litmus paper to measure the pH of solutions.  If the litmus paper turns red it indicates it is a very acidic solution.  If it turns blue, it indicates it is a very basic solution.

1. What is phenolphthalein used for?  If you add phenolphthalein to a solution and it turns pink, what does that tell you?  If it stays clear, what does that tell you?

Phenolphthalein is used to indicate whether a solution is acidic or basic.  If you add a drop of phenolphthalein to a solution and it turns pink, it is basic.  If you add a drop of phenolphthalein and it remains clear, it is acidic.

1. What is the formula we use to determine concentration of a solution?

Molarity = Mol

        L

1. If you mix 3.4 mol of HCl with 4.6 L of water, what is the concentration of the solution?  What is the pH?  pOH?

pH= 0.13   pOH =13.87

1. If you mix 1.48 mol of HF with 7.8 L of water, what is the concentration of the solution?  What is the pH?  pOH?

pH = 0.72  pOH = 13.28

1. If you mix 2.1 mol of NaOH with 5.6 L of water, what is the concentration of the solution?  What is the pH of the solution? pOH?

pOH = 0.43    pH = 13.57    (Remember, NaOH is a base.)